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The 2025 edition of the bestselling Pearson IIT Foundation has evolved into a blended resource, featuring a comprehensive coursebook for the students. The exemplary series will help students optimize their preparation with comprehensive and reliable content, video solutions, and interactive tests. The series is curated for students targeting esteemed examinations like JEE, NEET, NTSE, Olympiad, and other key competitive examinations. The new section containing practice questions for JEE and NEET aspirants, added in every chapter (Grades 8 to 10 only), gives that extra edge to the students in their exam preparations.

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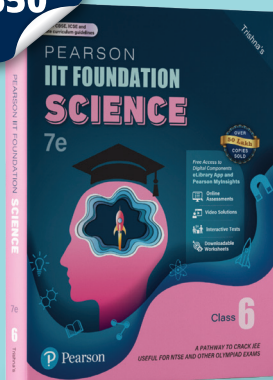
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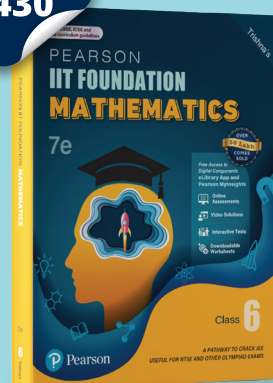
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- The Essence of Electricity
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- Mystery of Matter
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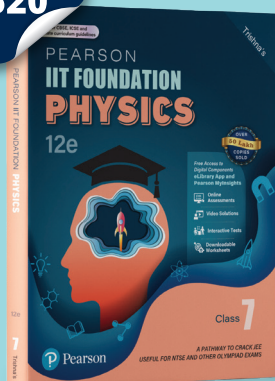


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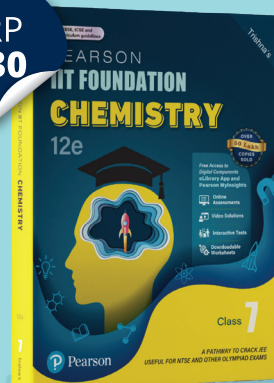
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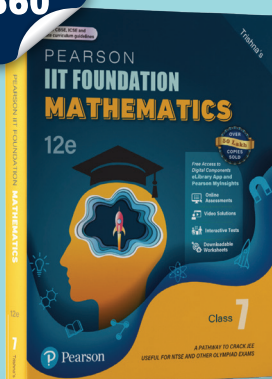
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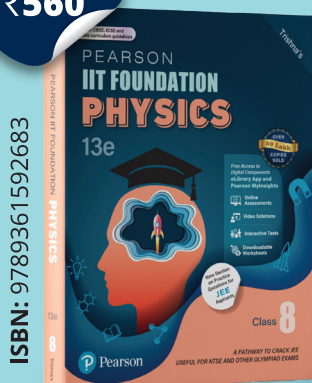


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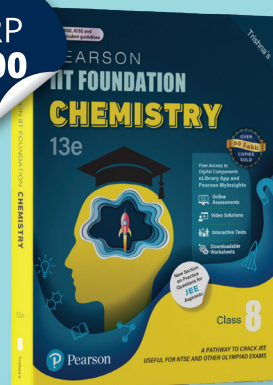
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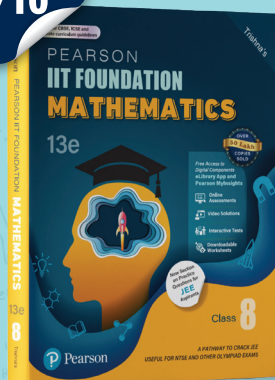
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- Linear Equations
- Sets
- Geometry
- Mensuration
- Statistics
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Chapter Insights ▼



Remember

Before beginning this chapter, you should be able to:

- understand the concept of matter and atoms.
- review the classification of matter and their

Remember section will help them memorise and review previous learnings on a particular topic

Key Points will help students identify essential points in a chapter



Key Ideas

After completing this chapter, you should be able to:

- understand the kinetic molecular theory and
- inter-relationship among various measurable properties of gases.
- understand mole as the chemical counting

Scientific Crypt

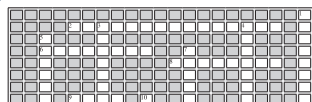
A tribute to Amedeo Avogadro (1776–1856)

Avogadro is the name familiar among chemists as the name associated with Avogadro number. Though, Avogadro was not the person responsible for the origin of the concept of Avogadro number, in recognition of his earlier contribution to chemistry, the name Avogadro number had been given at the time of introduction of concept of mole which was a major milestone in the advancement of chemistry.

Amedeo Avogadro was an Italian scientist born in Turin, Italy on 9th August 1776. As a son of distinguished lawyer, he had an established legal career. His interest in mathematics and sciences led him to spend most of his career as chair of Mathematical Physics at the University of Turin.

Avogadro's contribution to the field of chemistry should be understood in the context of the ideas prevailing at those times. Avogadro had the distinction of being the first scientist to bring out a clear distinction between atom and molecule. Though Avogadro had many publications in the field of chemistry his major contribution to the field was building on the work of chemist Joseph Louis Gay-Lussac which resulted in Avogadro's hypothesis. The studies made by Avogadro remained dormant during his lifetime. Later, it was after the work of Cannizzaro to introduce atomic weights based on Avogadro's work, that his hypothesis was brought into light and became Avogadro's law. As a person who laid the foundation and opened new vistas to quantitative chemistry, he is very well remembered, even after his demise as early as in July 1856.

CROSSWORD



FUN TIME

Scientific Crypt provides information about well-known scientists related to the chapter. The engaging activity is also given as crosswords to keep learners interested

Introduction section helps lay the foundation of the following chapter

INTRODUCTION

The study of matter in different perceptions is dealt by a branch of science named Chemistry. Matter can be classified on the basis of molecular composition as elements, compounds, and mixtures. Various elements are given names and symbols, and the names and formulae of various compounds have been derived based on certain rules. The study of naming these compounds and derivation of formulae for these are the prerequisites for representing chemical reactions in the form of chemical equations. Therefore, the language of chemistry,



Symbols and Formulae

All elements are represented by means of their symbols. A symbol represents the atom of an element, which may consist of a single letter or a combination of letters derived either from its English name or its Latin name.

A *symbol* is the shorthand notation that represents a single atom of an element.

Examples:

Element

Hydrogen

Mercury (formerly known as hydrargyrum)

Symbol

H

Hg

Concepts are explained in a well-structured and lucid manner

Konnect boxes offer add-on information of related topic

Konnect

A common example of neutralization is our body's response to heartburn.

Antacids are bases that neutralize the excess stomach acid.

Info Bytes!

The kinetic theory of gases is based on the assumption that gases are composed of a large number of tiny particles (atoms or molecules) that are in constant, random motion. This movement is what gives rise to pressure and temperature.

Info Bytes are interesting trivia and facts that engage young minds and keep the grey cells ticking!

Examples are given in the form of in-text exercises to reinforce the concepts learnt

Illustrative examples solved in a logical and stepwise manner

EXAMPLE 1.5

Gas laws are universally applicable for all gases whereas such universal laws could not be established for solids and liquids. Comment on this statement.

SOLUTION

Gases differ from solids and liquids in the extent of intermolecular forces of attraction. Since gases are characterized by negligible forces of attraction, the molecules behave independent of the neighboring molecules. Due to large intermolecular spaces, the coefficient of volume expansion is same for all gases. Moreover, in case of gases, the molecules are considered as point masses and hence, the volume occupied by the molecules is negligible in comparison to the total volume. Therefore, the physical behavior of all gases being similar, certain laws could be established which are universally applicable. In case of liquids and solids, the molecules are considered as rigid spheres and the coefficient of volume expansion is not uniform for all substances, such universal laws cannot be established.

EXAMPLE 1.6

Boyle's law says that pressure and volume are inversely proportional to each other. However, when a balloon is blown, both volume and pressure increase continuously. Justify.

SOLUTION

When a balloon is blown continuously, more and more air is forced into the balloon that increases the volume as well as the

POINTS TO REMEMBER

- A symbol is the shorthand notation that represents a single atom of an element.
- A formula is the shorthand notation which represents a single molecule of an element or compound.
- The properties of solids and liquids differ widely from substance to substance, but the gases generally obey some common rules known as the gas laws.
- Boyle's Law states that the volume of a given mass of a gas is inversely proportional to the pressure exerted by the gas at constant temperature.
- Vapor density (VD) of a gas or vapor is the ratio of the mass of a certain volume of gas or vapor to the mass of the same volume of hydrogen gas when their volumes are measured under similar conditions of temperature and pressure.
- A mole is defined as the quantity of substance that contains the same number of elementary particles or chemical units as the number of atoms present in 12 g of C-12 isotope.
- Dalton's law states that 'the total pressure exerted by a mixture of non-reacting gases taken in a container at a given temperature is the sum of the practical pressures of the

Points to Remember summarise concepts and topics presented in each chapter

PRACTICE QUESTIONS CLASSROOM WING

Very Short Answer Type Questions

Directions for questions from 1 to 3:
Fill in the blanks.

1. The percentage of sulphur in SO_2 is _____ and _____ in SO_3 .
2. The vapor density of SO_2 gas is _____ the vapor density of O_2 gas.
3. The molecular formula of a gas with vapor density 15 _____.

Short Answer Type Questions

26. A spherical balloon of volume 253 mL is to be filled with helium gas at STP from a cylinder containing gas at 20 atm and 127°C . If the cylinder can hold water, then calculate the number of balloons that can be filled up.

27. The composition of a gaseous mixture is $\text{N}_2 = 56\%$, $\text{O}_2 = 33\%$, $\text{Ar} = 11\%$ by mass. What is the partial pressure of nitrogen at 1 atm pressure?

(A) Reaction between an acid and a base from salt and water

(C) Preparation of oxygen from potassium nitrate
(D) Preparation of CO_2 from its constituents

9. What is the molarity of HCl solution containing 0.4 moles in 200 mL of the solution?
10. Calculate the volume occupied by 11 g of CO_2 gas at _____.

Essay Answer Type Questions

41. The density of a mixture of He and O_2 at STP is 1.25 g/L. What is the partial pressure of O_2 ?
42. A bottle is heated from 27°C to 127°C at constant pressure. What fraction of air gets expelled?
43. An oxide of Cr contains 68.5% of Cr by weight. Find _____.

Different typologies of questions have been included in the *Classroom Wing* which help students develop problem-solving skills

Competition Wing at the end chapters incorporated for competitive examination in-line with JEE syllabus

Graded as per complexity into Levels 1, 2 and 3.

PRACTICE QUESTIONS COMPETITION WING

Level 1: Apply your Concepts

Direction for questions from 1 to 8:

State whether the following statements are true or false.

1. Doubling the pressure of a gas at constant temperature, doubles the volume occupied by the gas.
2. _____.

16. 'ous' acid of a non-metal 'X' has 2 'O' atoms. Per acid of the same non-metal has the formula _____.

17. The ratio of the gram atomic weight of nitrogen and oxygen is _____.

18. Pressure exerted by water vapor in moist gas is called _____.

Level 2: Review Your Concepts

51. Out of oxygen, carbon dioxide, sulphur dioxide and nitrogen, which among these gases possesses maximum density at a given temperature and pressure?

52. If the Avogadro number is 9×10^{23} , calculate the weight of 1 mole of CO_2 gas.

53. A gas cylinder is filled with helium at 2000 mm. Due to _____.

Level 3: Assess Your Concepts

74. At constant temperature and at 4 atm pressure, the given amount of gas occupies 6 L. If the pressure is increased by 6 atm, calculate the percentage change in the volume of the gas.

75. A gaseous mixture contains 48 g of SO_2 and 33 g of CO_2 at constant temperature and pressure. Calculate the average molecular mass of the gaseous mixture.

76. Calculate the ratio of rates of diffusion of three isotopes of hydrogen protium, deuterium and tritium, respectively, at the same temperature and pressure.

26. A spherical balloon of volume 253 mL is to be filled with helium gas at STP from a cylinder containing gas at 20 atm and 127°C . If the cylinder can hold 2 L of water, then calculate the number of balloons that can be filled up.

27. The composition of a gaseous mixture is $\text{N}_2 = 56\%$, $\text{O}_2 = 33\%$, $\text{Ar} = 11\%$ by mass. What is the partial pressure of nitrogen at 1 atm pressure?

28. A gold manufacturing company manufactures gold coins with absolutely pure gold. If the mass of each gold coin is 9.85 g and 1000 gold coins are manufactured per day, calculate the number of moles of gold consumed per day. (At. wt. of gold is 197 amu)

29. A closed cylinder is first evacuated and then connected to another cylinder of 75 L volume containing an ideal gas 'X'. The pressure of the gas in the second cylinder _____.

33. What is the volume of 50% (w/v) H_2SO_4 required for the liberation of 5.6 L of hydrogen gas at STP on its reaction with magnesium?

34. Calculate the percentage composition of glucose.

35. Empirical formula of a compound is $\text{C}_2\text{H}_5\text{N}$. If the empirical formula weight is equal to one-fourth of its _____.

52. If the Avogadro number is 9×10^{23} , calculate the weight of 1 mole of CO_2 gas.

53. A gas cylinder is filled with helium at 2000 mm. Due to leakage, the pressure dropped to 1500 torr in 40 min. When the same cylinder is filled with another gas at the same pressure, the pressure dropped from 2000 mm to 1500 mm in 200 min. What is the molecular weight of the gas?

54. A cation A^{2+} and an anion B^{2-} combine to form a compound. Similarly, C^{2+} and D^{2-} also combine to form a compound. When the molten compounds are subjected to electrolysis, the gaseous products obtained _____.

Icons are used for categorisation of questions in Classroom Wing and Competition Wing

PRACTICE QUESTIONS JEE SPECIAL WING

Multiple-Answer Type Questions

Directions for Questions 1 to 5: Each question has **FOUR options (a), (b), (c), and (d). ONE OR MORE THAN ONE** of these four option(s) is/are the correct answer(s).

- The constituents of biogas are _____.
(a) Methane (b) Carbon dioxide
(c) Hydrogen (d) Nitrogen
- The value of solar constant is _____.
(a) $1.4 \text{ kJ m}^{-2} \text{ s}^{-1}$ (b) $4.1 \text{ kJ m}^{-2} \text{ s}^{-1}$
(c) 1.4 kW m^{-2} (d) $1.4 \text{ kW m}^{-2} \text{ s}^{-1}$
- Which of the following sources of energy can provide energy over a long period of time?
(a) Sun (b) Wind
(c) Coal (d) Petrol
- Which of the following types of coal contains the highest percentage of carbon?
(a) Peat (b) Lignite
(c) Bituminous (d) Anthracite
- The energy received on the periphery of the atmosphere of an area of 75 m^2 in one hour is ____.

JEE-NEET Special Wing is curated specifically for targeted preparation of JEE and NEET examinations

Hints and Explanations

CLASSROOM WING

Very Short Answer Type Questions

- Non-renewable
- Renewable
- Ignition temperature
- Geothermal
- Nuclear fusion
- Calorific
- Hydrocarbons
- Peat
- Infrared radiations

- Petrol – presence of hydrogen.
- A device that converts wind energy to mechanical or electrical energy.
- A device which directly converts solar energy into electrical energy.

Short Answer Type Questions

Note: This is how you should frame your answer covering.

- Wind mill, turbine.
- (i) Solid fuels
(ii) Liquid fuels
(iii) Gaseous fuels

Hints and Explanations for key questions of classroom wing, competition wing and JEE-NEET special wing along with highlights on the common mistakes that students usually make in the examinations

Competition Wing

Level 1: Apply Your Concepts

Explanation for questions from 1 to 5:

- Calorific value.
- Hydrocarbons
- Ignition temperature
- Solar energy.
- The minimum temperature at which the substance begins to burn.

Explanation for questions from 6 to 10:

- False. Ignition temperature is the minimum temperature at which a substance begins to burn, not melt.
- True.
- False. Tidal energy is renewable as it relies on the gravitational effects of the moon and sun, which are repetitive and consistent.
- True.
- False. This description is for nuclear fusion.

Explanation for questions from 21 to 27:

- (b). Nuclear fission involves the splitting of a heavy nucleus, while nuclear fusion involves the combination of lighter nuclei.

Explanation for questions from 36 to 44:

- Silver is the metal which is used for connecting wires that connect solar cells in a solar panel, due to its large conductivity.
- The most suitable fuel burns completely and does not release (expel) any carbon particles with smoke.
- Solar energy provides wind energy, light energy and heat energy.
- In case of box-type solar cooker, more energy will be trapped by using a box, whose inner surfaces are painted black and which cannot conduct more energy to the surroundings. Now from the given options, if we choose a double wall plastic box, whose inner surfaces are painted black, then the efficiency will be more.
- The net output = output – input = $1000 - 200 = 800 \text{ J}$.
- A good source of energy should have large, net energy output. Among the given options, a source of energy which has input energy of 100 J and an output energy of $12,500 \text{ J}$ is the best source of energy, because the net energy output is $(12,500 - 100) \text{ J} = 12,400 \text{ J}$.
- More is the velocity of wind, more will be the kinetic energy of blades of a windmill due to which more amount of electrical energy can be produced. That is, the output of a windmill increases with increase in the



WORKSHEET CHAPTER 2: DYNAMICS

DW

DW: Downloadable worksheets supplementing textual exercises for hand-on learning

Directions for questions from 1 to 5:

Select the correct alternative from the following options.

1. A bullet of mass 100 g moving with a velocity of 100 m s^{-1} is stopped by a plank. If the bullet travels a distance of 10 cm in the plank before it comes to rest, find the average force exerted by the plank.
(a) 100 N (b) 1000 N
(c) 2500 N (d) 5000 N
2. Find the power developed by an engine in filling water

a rough surface, the total work done by the frictional force is zero.

8. The orbital velocity of a satellite is inversely proportional to the radius of the orbit.
9. When a stone is tied to a string and rotated in a circular path, the work done by the tension of the string is zero.
10. A satellite revolving in an orbit closer to the surface of the earth has more speed than one which is moving in an orbit of higher radius.

AT

AT: Attemptable Tests for assessing understanding of concepts and getting a clear picture of your exam readiness

AI Assessments

Which of the following is larger unit of density?

- ☐ g m^{-3}
- ☐ g km^{-3}
- ☐ kg m^{-3}
- ☒ g cm^{-3}

Questions 1 of 4

Next

19. A standard vernier callipers with positive zero error of 0.03 cm is used to measure the diameter of a sphere. If the MSR is 5.5 cm and the diameter observed is 5.56 cm, then what is the true VCD?
20. The main scale of a vernier callipers has 10 divisions for each 1 cm and 10 divisions on its vernier scale such that $N \text{ VSD} = (N-1) \text{ MSD}$. What is the least count of the vernier callipers?
21. If the least count of a screw gauge is 0.001 mm and if the screw advances by 0.5 mm for its two complete rotations, how many divisions does the circular scale of



Video icon: Audio and visual learning to difficult questions for effective learning and retention

IIT Foundation Series Grade IX

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- Wave Motion and Sound

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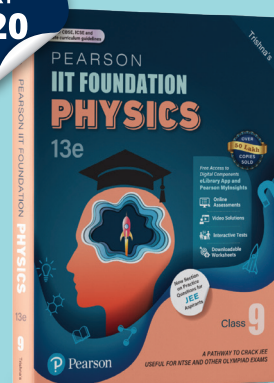
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- Magnetism-I

LIGHT

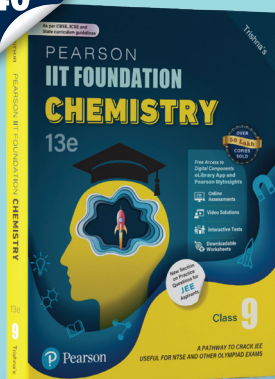
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- Atomic Structure
- Chemical Kinetics and Chemical Equilibrium

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- Periodic Classification of Elements
- Chemical Bonding
- Water
- Metal and Non-Metals

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- Simple Interest and Compound Interest

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- Logarithm

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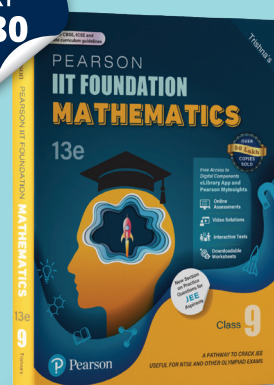
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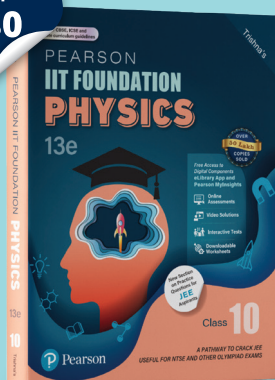


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- Electrochemistry
- Periodic Table
- Chemical Bonding

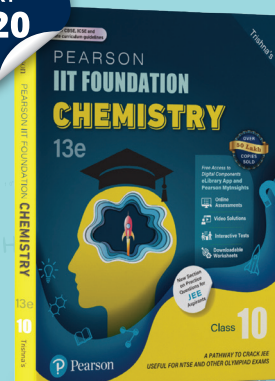
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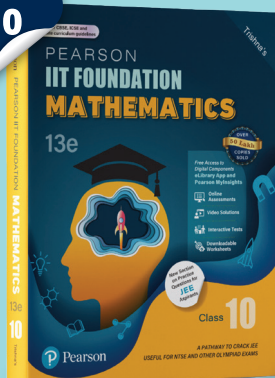
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- Pair of Linear Equations in Two Variables
- Quadratic Equations
- Progressions

GEOMETRY

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- Mensuration
- Coordinate Geometry

TRIGONOMETRY

- Trigonometry

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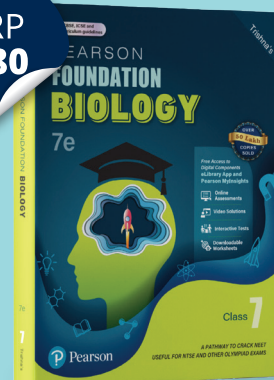
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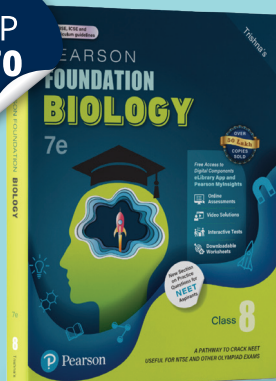
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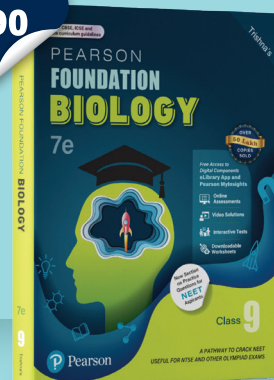
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- Tissues
- Diversity in Living World - plants
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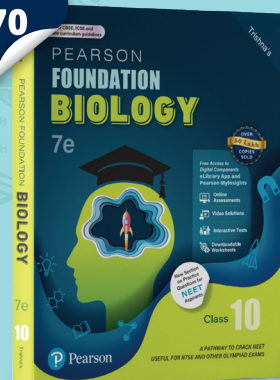
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- Cell: The Basic Unit of Life
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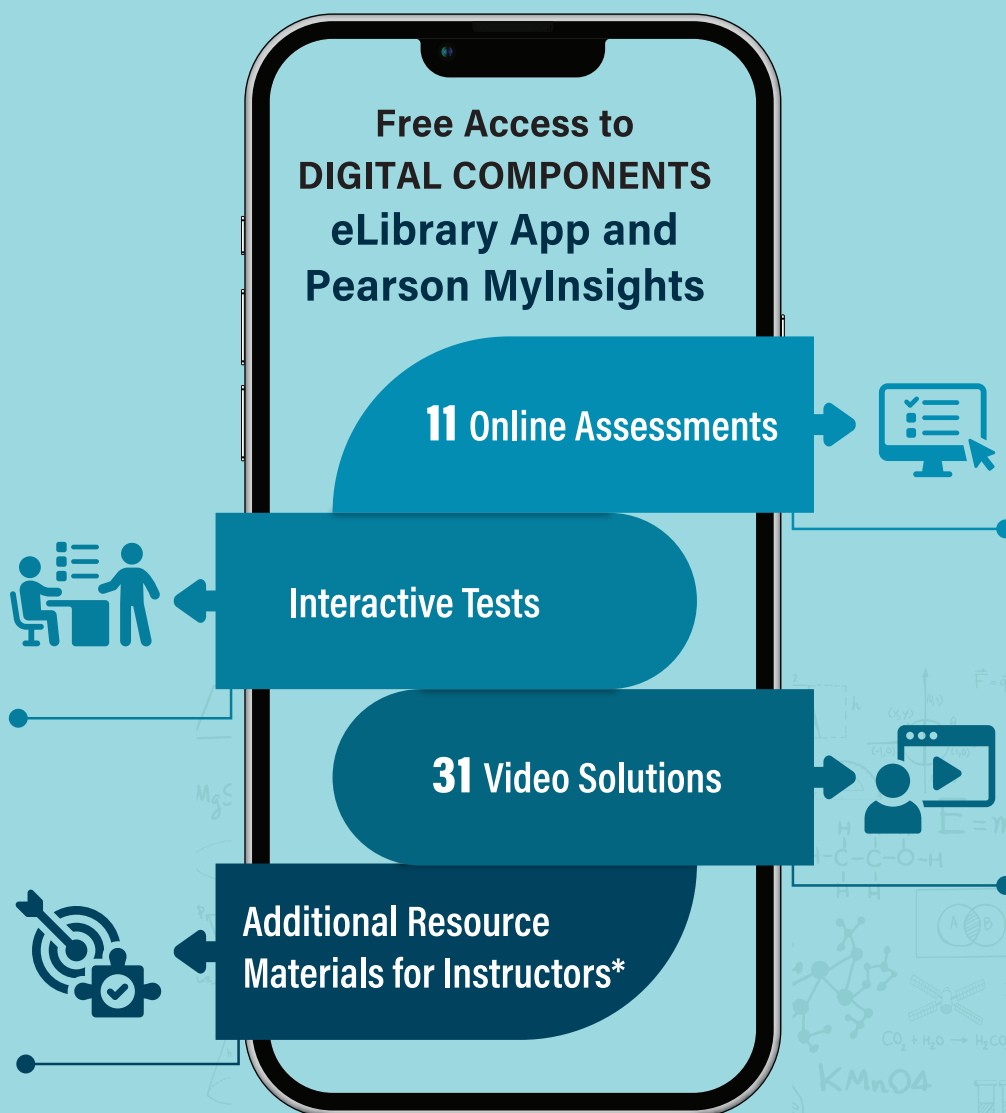
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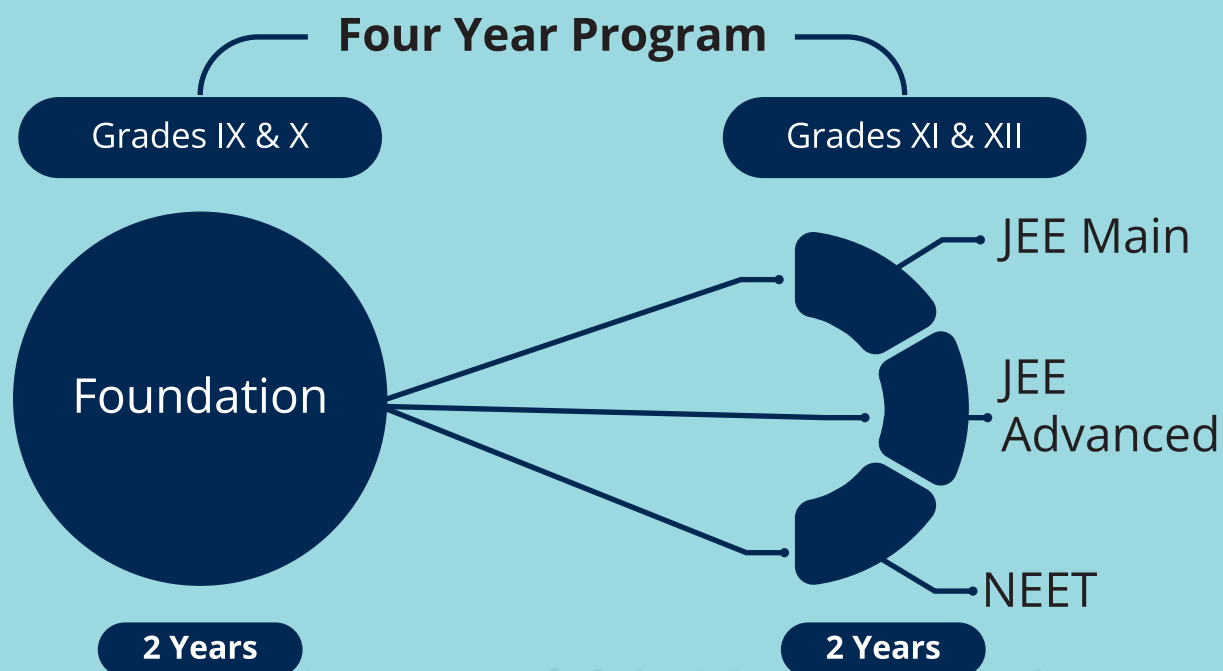
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